Date 12 Aug 2021

Full chapter at a glance

Stoichiometry

i No of moles

$$n = \frac{\text{mass}}{\text{molar mass}}$$

ii Mass from no of moles

$$\text{Mass} = n \times \text{Molar mass}$$

iii Conversion relationship

$$\frac{V}{22.4\,dm} = n = \frac{\text{Mass}}{\text{Molar mass}}$$

iv Total No of identities:

Total no of atoms = n x Na x No of Particles in one molecule/Compound

v Total No of identities

Total no of electrons = n x Na x No of electrons in one molecule

vi Total No of identities

Total charge on any ion = n x Na x No of Protons in one molecule

vii Total No of identities

Total charge on any ion = n x Na x Charge on one ion x 1.6×10^{-19} C

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viii amu to gram

$1\ amu = 1.661 \times 10^{-24}\ g$

ix gram to amu

$1g = \frac{1}{1.661 \times 10^{-24}}\ amu$

$1g = 6.023 \times 10^{+23}$

x Percentage Composition

$$\frac{\text{Total mass of element in Compound}}{\text{Total mass of compound}} \times 100$$

Al. $$\frac{\text{Mass of element} \times \text{No of atoms in Compound}}{\text{Molar mass of the Compound}} \times 100$$

xi % Yield

$$\frac{\text{Actual Yield}}{\text{Theoratical Yield}} \times 100$$

"Remember me in your Precious Prayers...★"

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